

T.

for the α liquid phase:

$$\ln \gamma_H^\alpha = \frac{8163 (.9098)^2}{(8.314)(348)}$$

$$\gamma_H^\alpha = 10.3331$$

$$\ln \gamma_E^\alpha = \frac{8163 (.0902)^2}{(8.314)(348)}$$

$$\gamma_E^\alpha = 1.0236$$

for the gas phase, assuming $\phi_i = 1$

$$P_H = y_H P = \gamma_H x_H P_H^s$$

$$P_E = y_E P = \gamma_E x_E P_E^s$$

$$P_H = (10.3331)(.0902)(1.202) = 1.120 \text{ atm}$$

$$P_E = (1.0236)(.9098)(0.8714) = 0.8115 \text{ atm}$$

b) $P = P_H + P_E = 1.932 \text{ atm}$

c) $y_H = 1.120 / 1.932 = 0.5799$
 $y_E = 0.8115 / 1.932 = 0.4201$

for β liquid phase:

$$(\gamma_E x_E)^\alpha = (\gamma_E x_E)^\beta$$

$$(1.0236)(.9098) = \left[\frac{8163 (1 - x_E^\beta)^2}{(8.314)(348)} \right] x_E^\beta$$

$$0.9313 = (2.821)(1 - 2x_E^\beta + (x_E^\beta)^2) x_E^\beta$$

$$0.9313 = \exp [2.821(1-x_E^B)^2] x_E^B$$

$$\ln(0.9313) = 2.821(1-x_E^B)^2 + \ln x_E^B$$

$$-.0712 = 2.821(1-x_E^B)^2 + \ln x_E^B$$

But, we know that from regular soln,
~~the~~ the liquids should be symmetric.

a) $\boxed{\therefore \text{try } x_E^B = 0.0902}$

$$-.0712 \stackrel{?}{=} 2.821(.9098)^2 + \ln(.0902)$$

$$-.0712 = -.0704 \quad \text{close enough}$$

II. (a)

at the azeotrope, assuming $\phi_{is} = 1$

$$\gamma_i x_i P_i^s = y_i P$$

$$\gamma_i = P / P_i^s$$

$$\gamma_1 = 1/.8018$$

$$\ln \gamma_1 = 0.22090$$

$$\gamma_2 = 1/.3430$$

$$\ln \gamma_2 = 1.07002$$

$$0.22090 = [A + 2(.67993B - A)(.57991)](.17647)$$

$$1.07002 = [B + 2(1.4707A - B)(.42009)](.33630)$$

~~0.22090~~

$$1.25177 = A - 1.15982A + 0.78860B$$

$$3.18174 = B - .84018B + 1.23565A$$

$$1.25177 = -0.15982A + 0.78860B$$

$$3.18174 = 1.23565A + 0.15982B$$

$$A = 2.3091$$

$$B = 2.0553$$

(b) at $x_1 = 0.2, x_2 = 0.8$

$$\ln \gamma_1 = [2.3091 + 2(1.3975 - 2.3091)(.14528)](.17305)$$

$$\ln \gamma_2 = [2.0553 + 2(3.396 - 2.0553)(.85471)](.02111)$$

$$\gamma_1 = 4.452$$

$$\gamma_2 = 1.0961$$

$$y_1 P = (4.452)(.2)(.8018) = .7139$$

$$y_2 P = (1.0961)(.8)(.3430) = .30077$$

$$P = 1.015$$

$$y_1 = 0.7036$$

$$y_2 = 0.2964$$